

## Practice Problems on Lewis Structures

1. Draw good Lewis structures for each of the following:

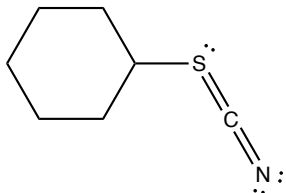
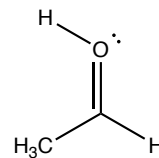
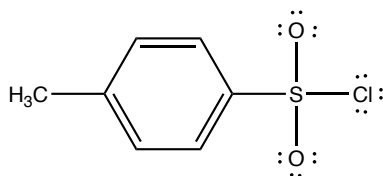
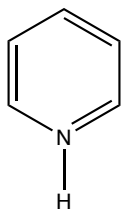
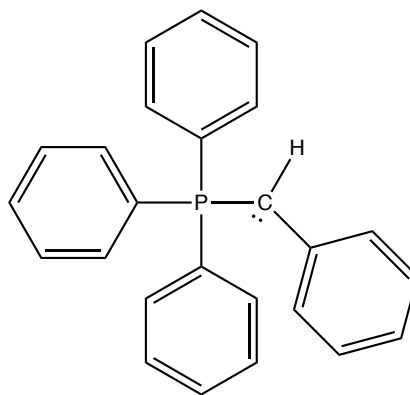
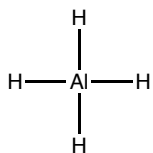
CO

NO<sub>2</sub>

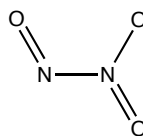
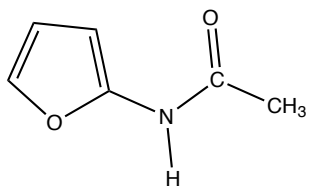
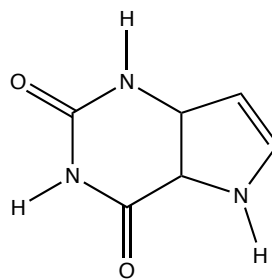
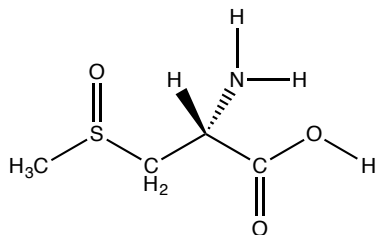
C<sub>2</sub>H<sub>6</sub>O (2 isomers)

BC<sub>3</sub>H<sub>9</sub> (4 isomers)

2. For each of the following determine the formal charges.



3. Draw lone pairs of electron where needed.



4. Write the hybridization of the atoms with an arrow.

