

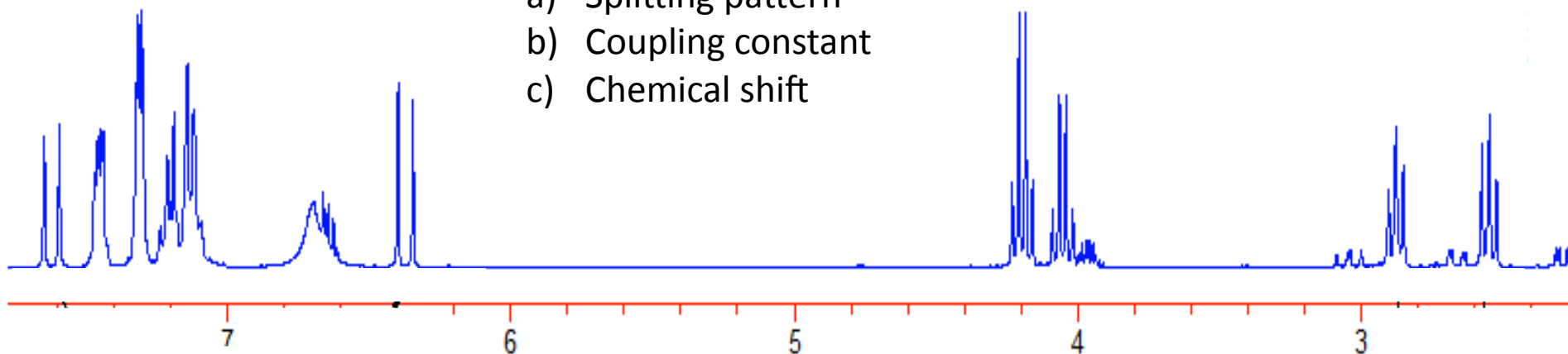
Assume that reaction went to 100% completion (i.e. no starting material remaining)

$^1\text{H-NMR}$  spectrum of crude reaction mixture shown below

Need to determine product ratio in crude reaction mixture using integration values

Assign signals using:

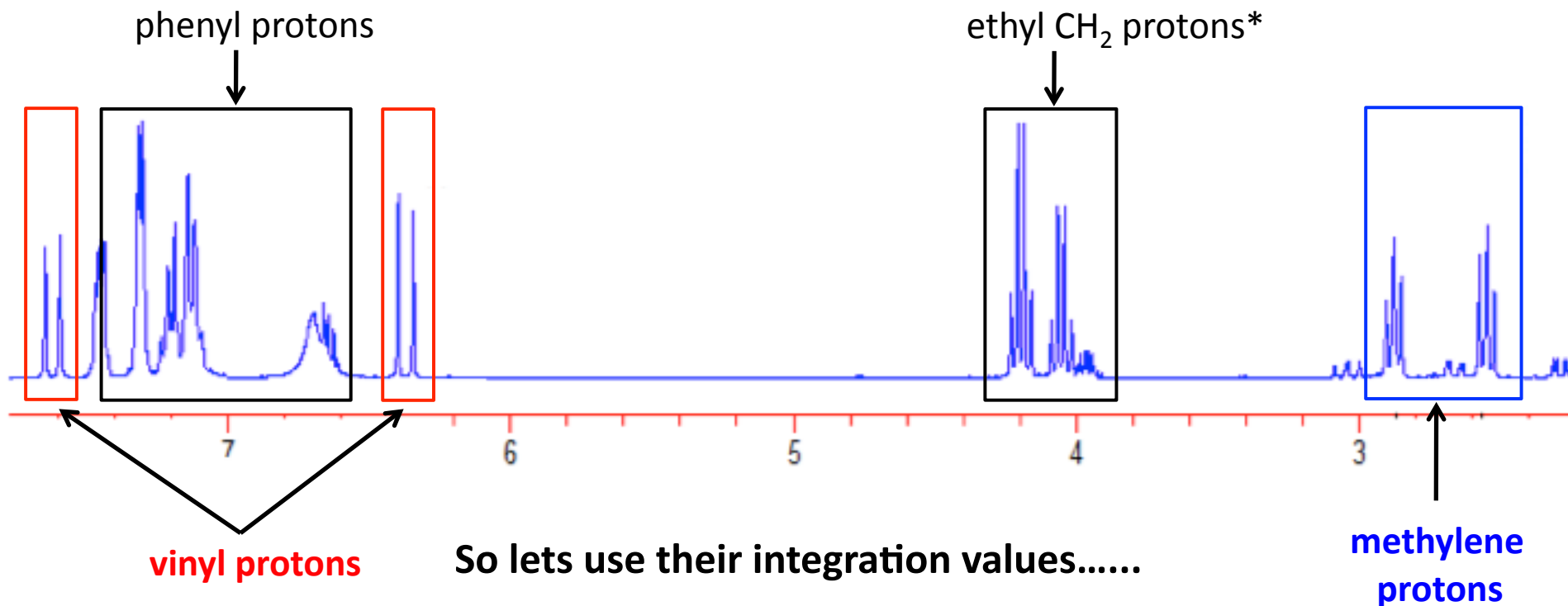
- Splitting pattern
- Coupling constant
- Chemical shift



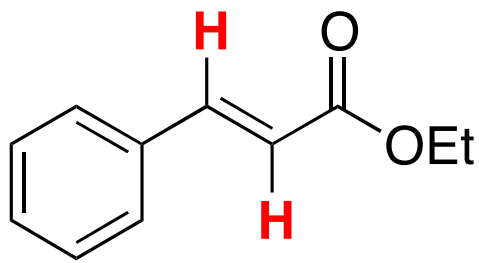
Which signals do we pick to integrate?

Ethyl and phenyl group signals don't help because they are overlapped or not well resolved

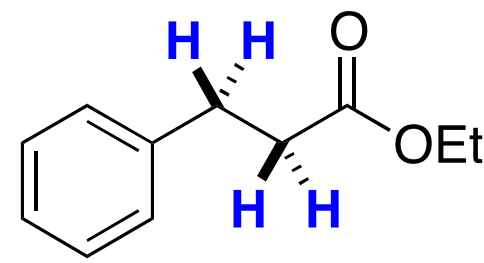
**Vinyl** and **methylene** signals are well separated and clear of other signals



So lets use their integration values.....



Product A



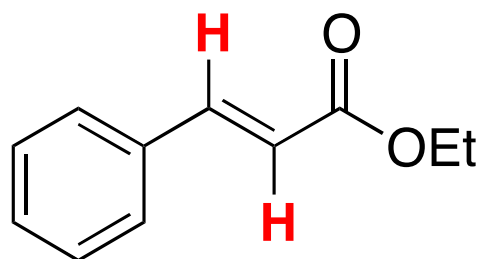
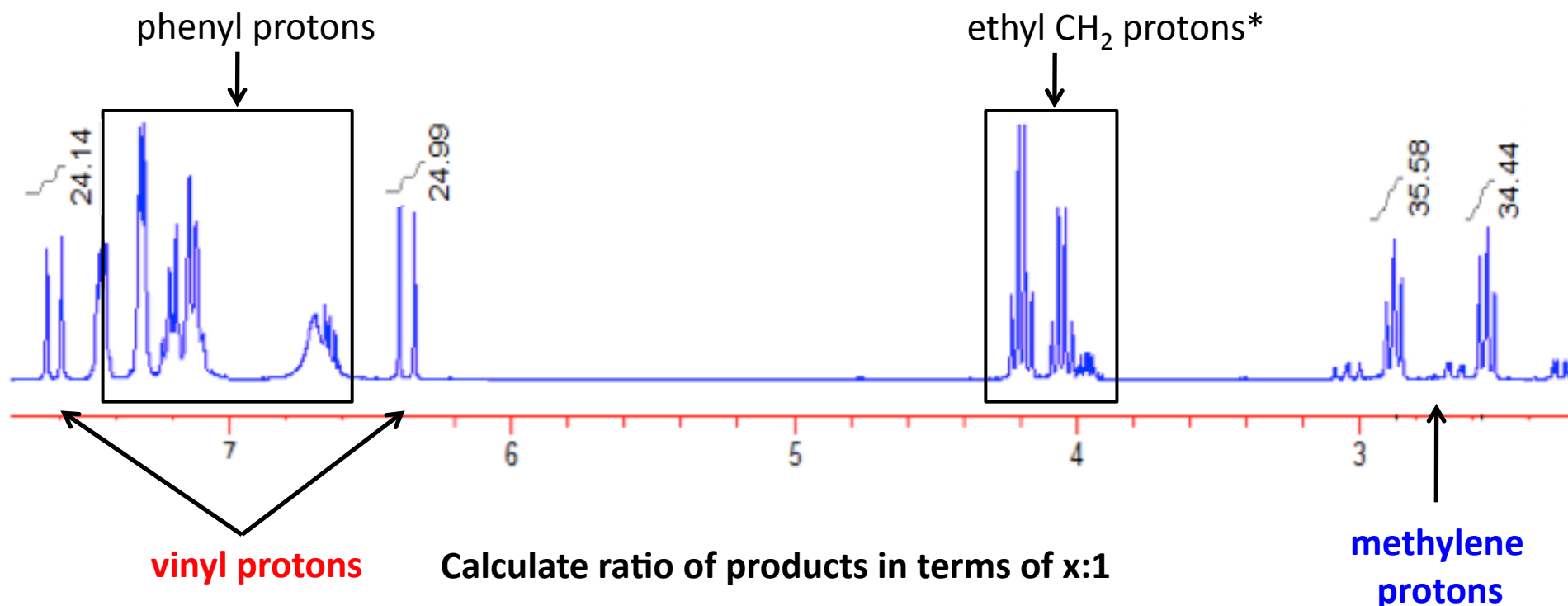
Product B

\*ethyl CH<sub>3</sub> protons overlapped ~1.4 ppm

**Vinyl** proton integrations are 24.99 and 24.14 (“raw” or “absolute” values)

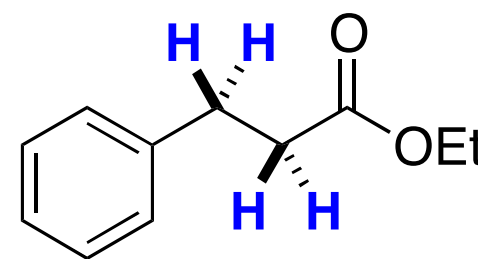
**Methylene** proton integrations are 35.58 and 34.44

There are 2 **methylene** protons per **vinyl** proton



**Product A**

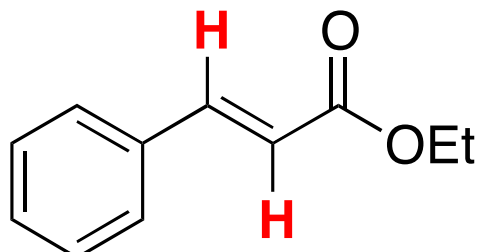
\*ethyl CH<sub>3</sub> protons overlapped ~1.4 ppm



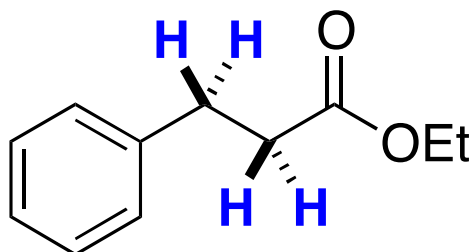
**Product B**

**Vinyl** proton integrations are **24.99** and **24.14**

**Methylene** proton integrations are **35.58** and **34.44**



Product A



Product B

$$\text{Ratio products} = \frac{\text{Integration 1 vinyl proton}}{\text{Integration 1 methylene proton}} = \frac{24.99}{34.44 \div 2} = 1.45 : 1$$

↑  
Why  $\div 2$  ?

because there are 2 **methylene** protons per **vinyl** proton

Could also use

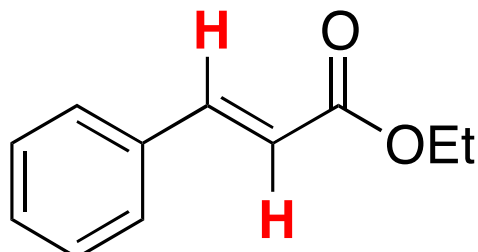
$$24.14 / (35.58 \div 2) = 1.36 : 1$$

$$24.99 / (35.58 \div 2) = 1.40 : 1$$

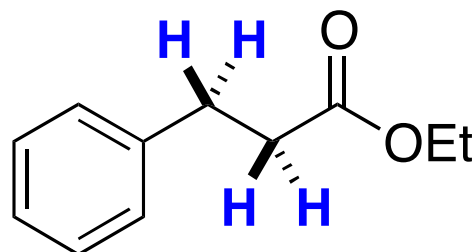
$$24.14 / (34.44 \div 2) = 1.40 : 1$$

Any of the above product ratios are OK to use

What can be concluded about the reaction we performed?



Product A



Product B

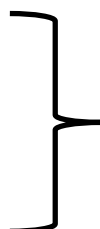
Product ratios from  $^1\text{H-NMR}$  signals

$$24.99 / (34.44 \div 2) = 1.45 : 1$$

$$24.14 / (35.58 \div 2) = 1.36 : 1$$

$$24.99 / (35.58 \div 2) = 1.40 : 1$$

$$24.14 / (34.44 \div 2) = 1.40 : 1$$



Ratio Product A : Product B =  $\sim 1.40 : 1$

- Reaction gave desired product (A) as major product
- Only slight excess of A thus selectivity of reaction is low

## Summary

- 1) Assign signals using splitting pattern,  $J$  value, and chemical shift
- 2) Look for signals that are well resolved and clear of other signals (i.e. not overlapped or cluttered)
- 3) Use the raw integration values of each signal that you pick
- 4) Account for the number of protons in the molecule that give each signal
- 5) Express ratio in terms of  $x:1$  ( $x$  = major product)