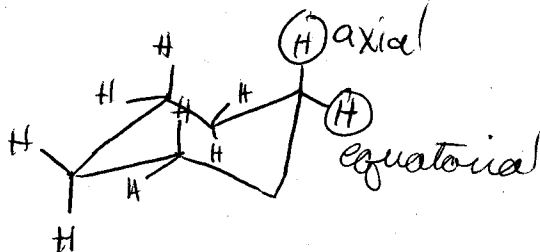
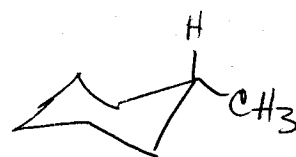
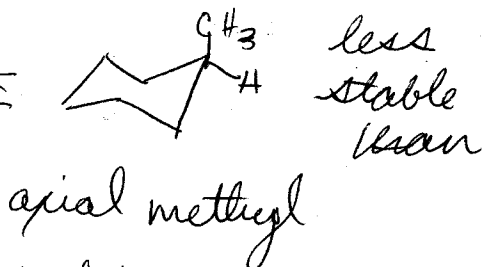
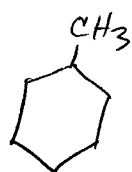


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Recall: Chair conformation of cyclohexane

w/a substituent

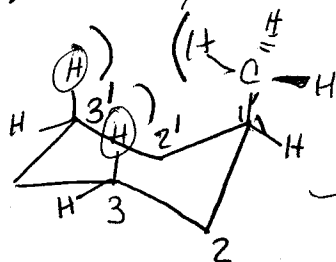
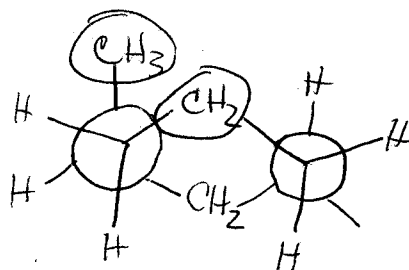


equatorial methyl

 $\sim 1.8 \text{ Kcal/mol}$

- Axial methyl conformation has higher internal energy because of internal steric repulsions

"1,3-diaxial interactions" • numbering indicates relative spacing

Newman
projection

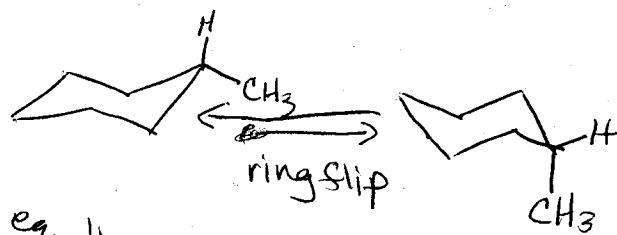
- like gauche butane - one of two 1,3-diaxial interactions

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- 1) Where is the other 1,3-diaxial interaction?
- 2) Draw the Newman projection for equatorial sub. methyl conf.

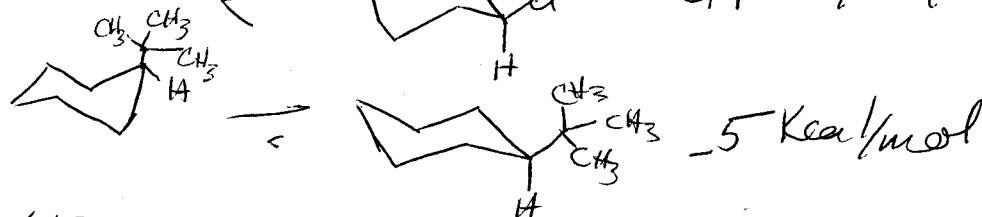
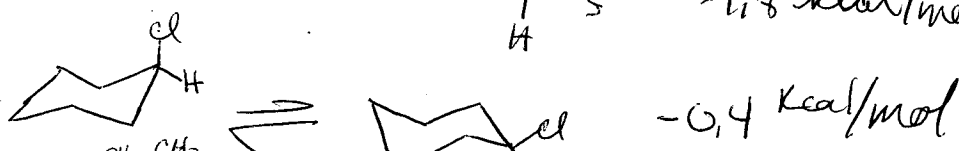
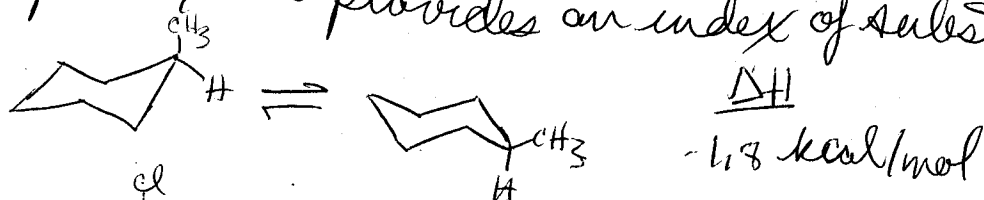
Cyclohexane has 2 chair conformations. If you convert from one to the other, each axial H goes to equatorial, and vice versa



Chair conf. interconversion

eq. lies to left b/c eq. is more stable

Note: position of conf. equilibrium for monosubstituted cyclohexanes provides an index of substituent size.



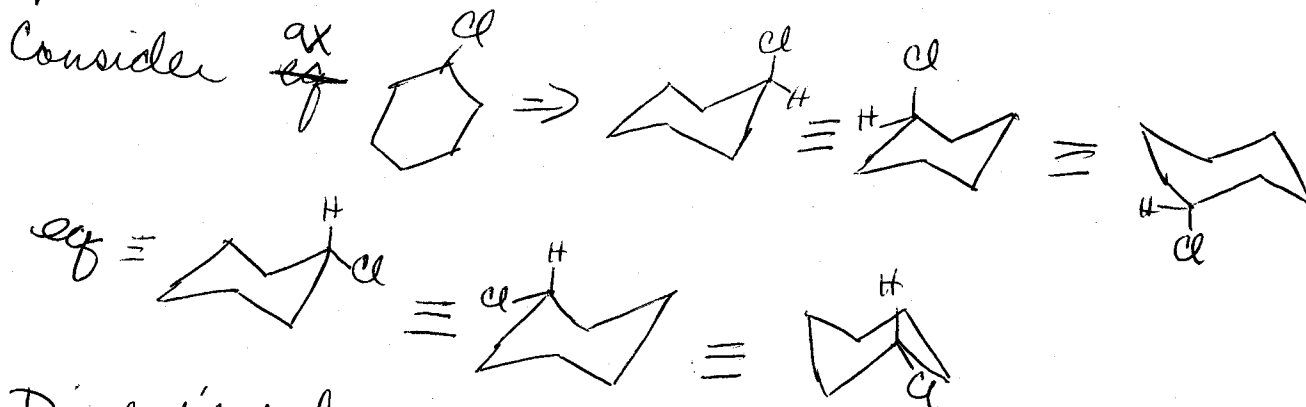
t-butyl is BIG!

t-Butyl cyclohexane is effectively "locked" into one of two chair conformations.

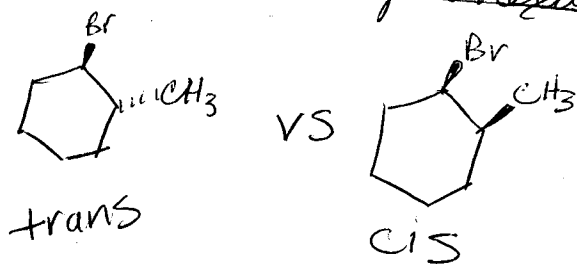
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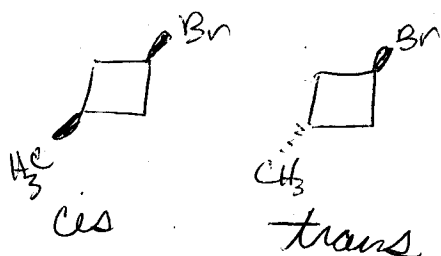
Note: Must be able to recognize when different drawings represent the same molecule.



Disubstituted cyclohexanes



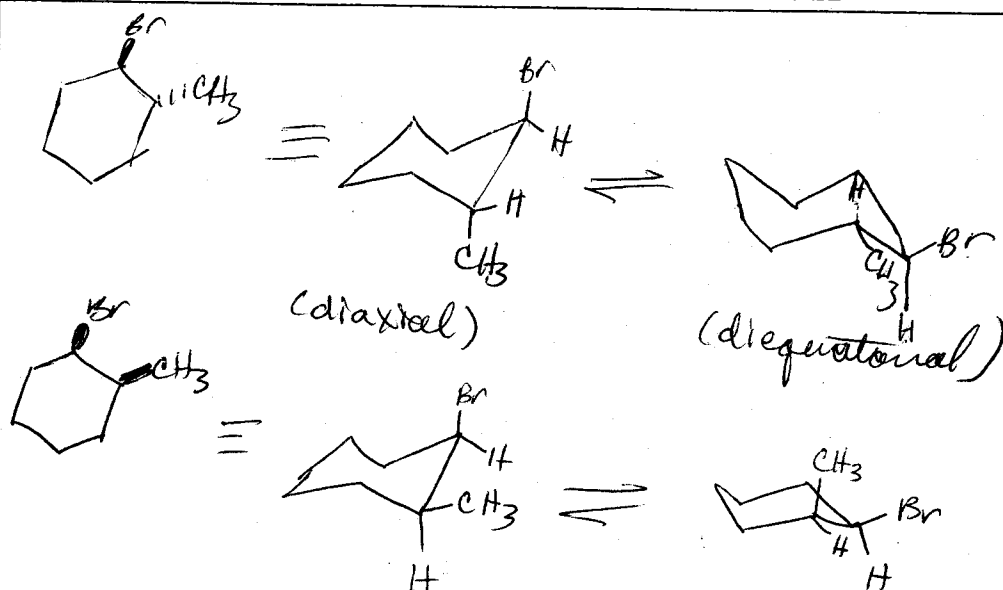
• This is general for rings



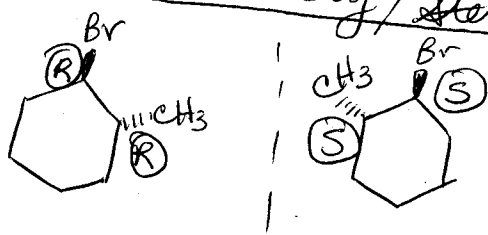
We must understand how cis/trans is manifested in chair conformations

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You fill in 1,3 + 1,4-disubstitution cases
Wedges & dashes do not correlate axial + equatorial!
Stereochemistry / Stereoisomerism

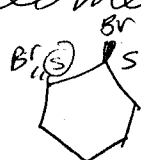
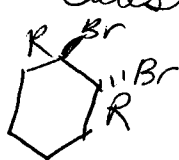


enantiomers

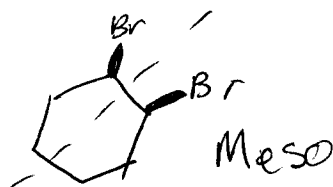
cis - another pair of enantiomers

cis vs trans

⇒ diastereomers

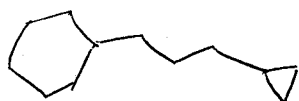


enantiomers



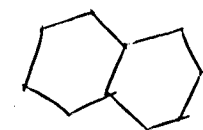
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Molecules w/ ≥ 2 rings, ...

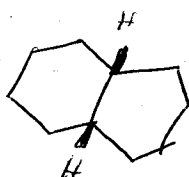


Share 1 atom or more.

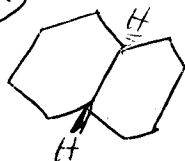
Share two atoms (i.e., 1 bond) $\Rightarrow$ "fused rings"



decalin



cis-decalin ("cis ring fusion")



trans decalin ("trans ring fusion")