

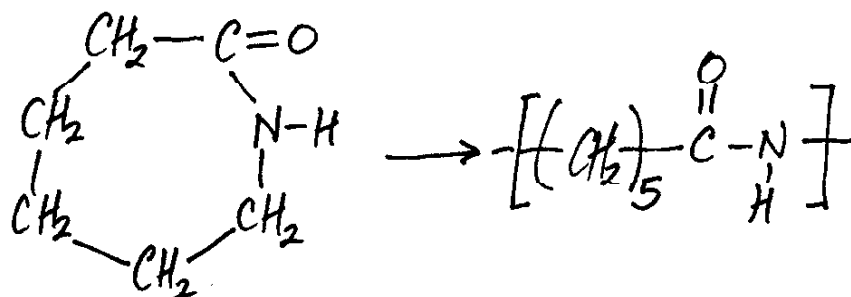
Chem 654 / Exam #1

Key

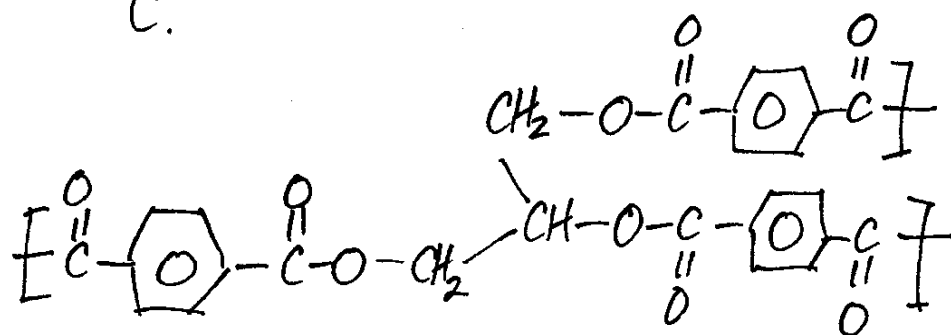
①



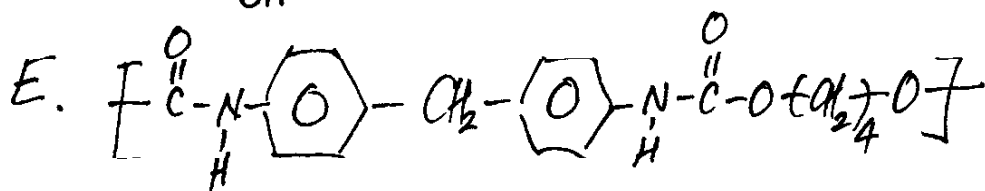
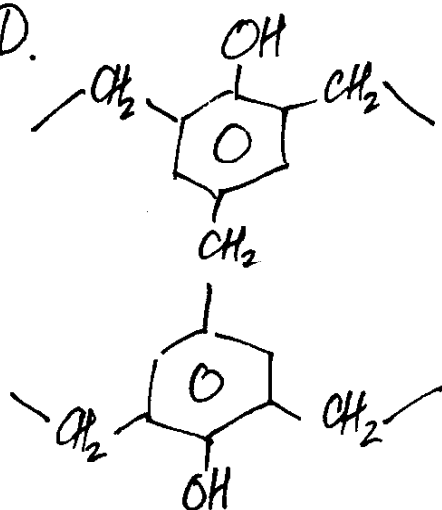
B.



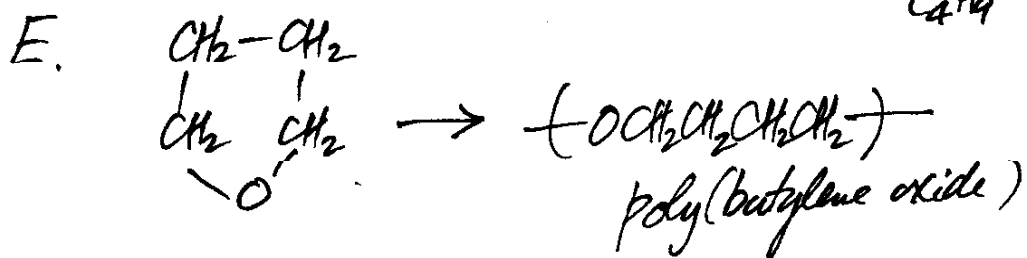
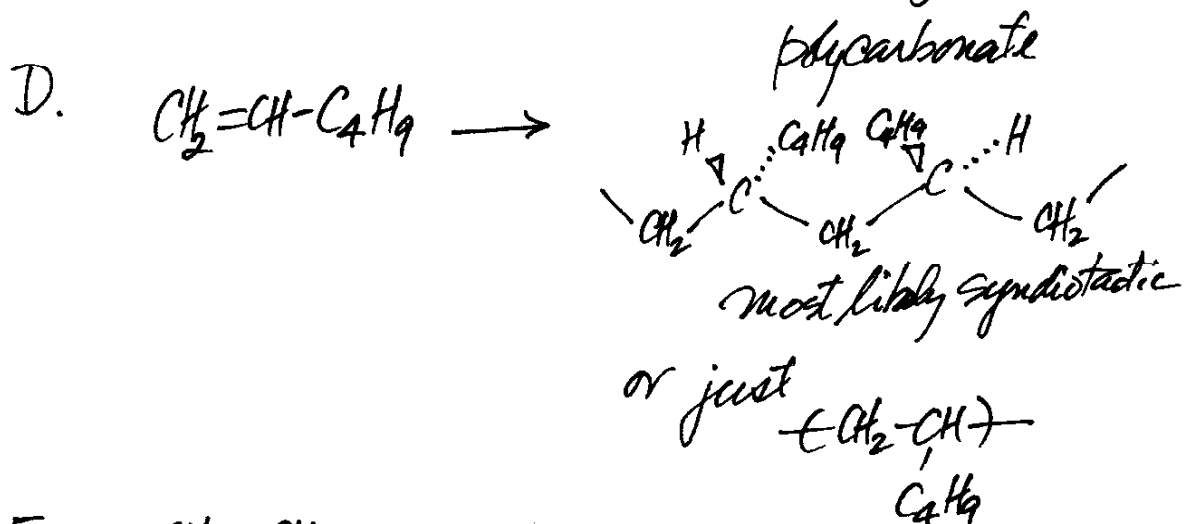
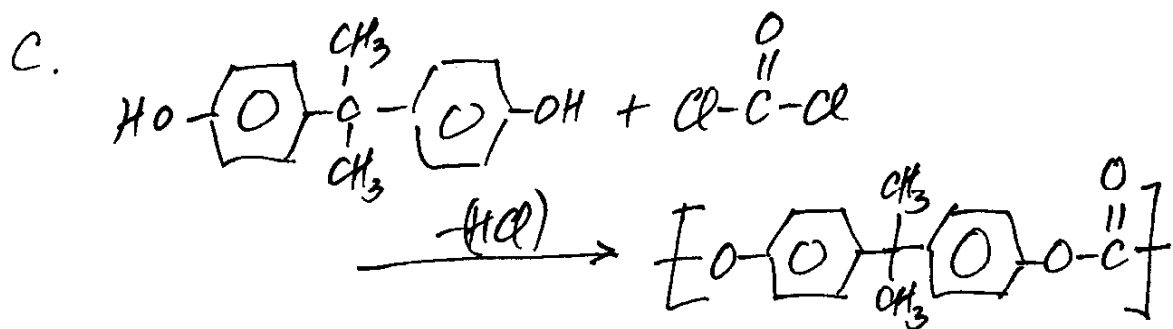
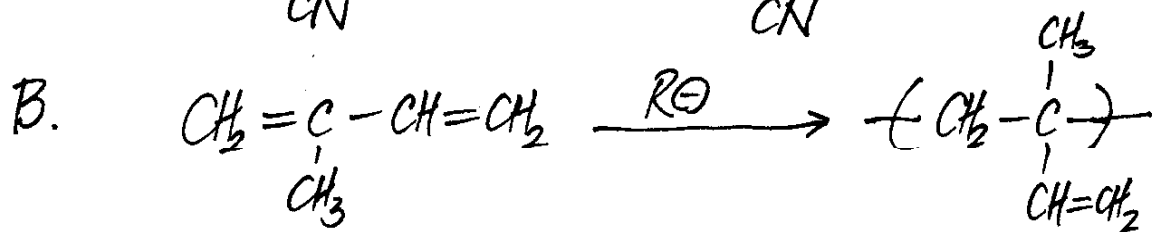
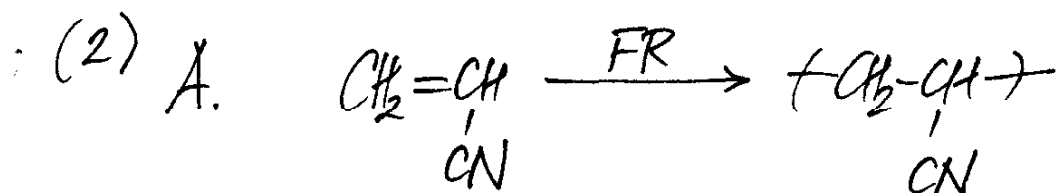
C.



D.

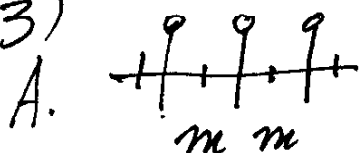


(2)

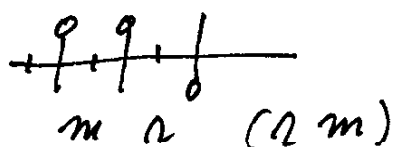


(3)

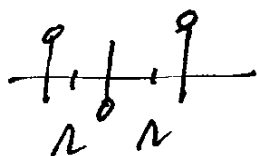
(3)



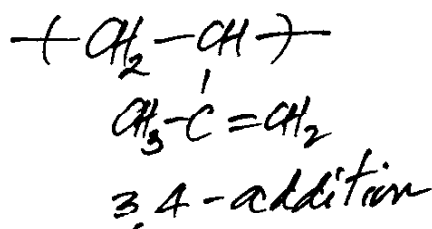
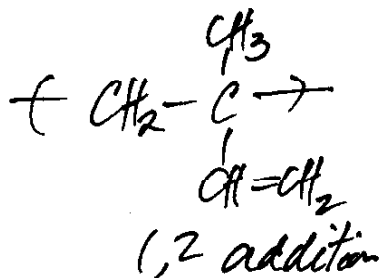
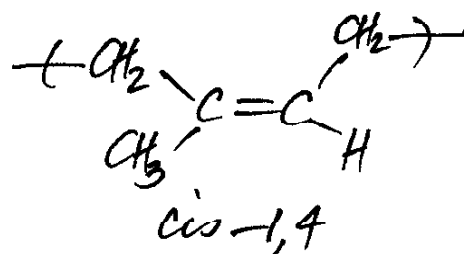
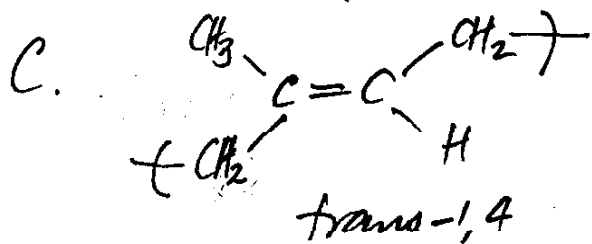
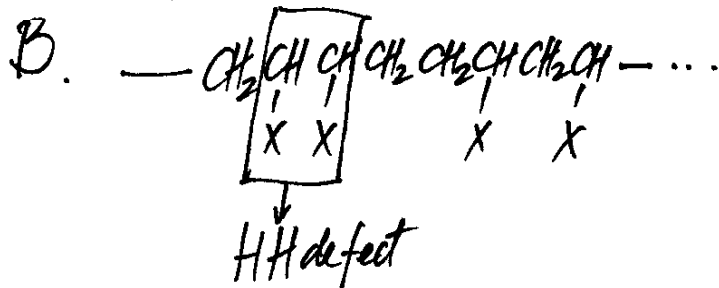
$$P(mm) = \frac{2!}{2!0!} p_m^2 = p_m^2$$



$$P(mn) = \frac{2!}{1!1!} p_m (1-p_m) = 2p_m (1-p_m)$$



$$P(nn) = \frac{2!}{0!2!} p_m^0 (1-p_m)^2 = (1-p_m)^2$$



D. In free radical polymerization,

$$\frac{d(R_x^\bullet)}{dt} = 0 \quad \text{for } 1 \leq x \leq \infty$$

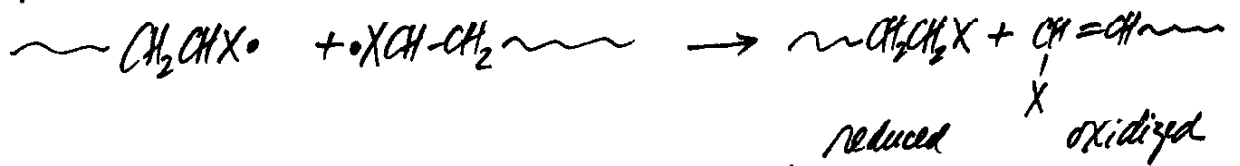
E. Polycondensation

$$X_x = p^{x-1} (1-p) \quad \text{where } p \text{ is the extent of reaction is indep of chain length, i.e., the same } p \text{ for } 1 \leq x \leq \infty$$

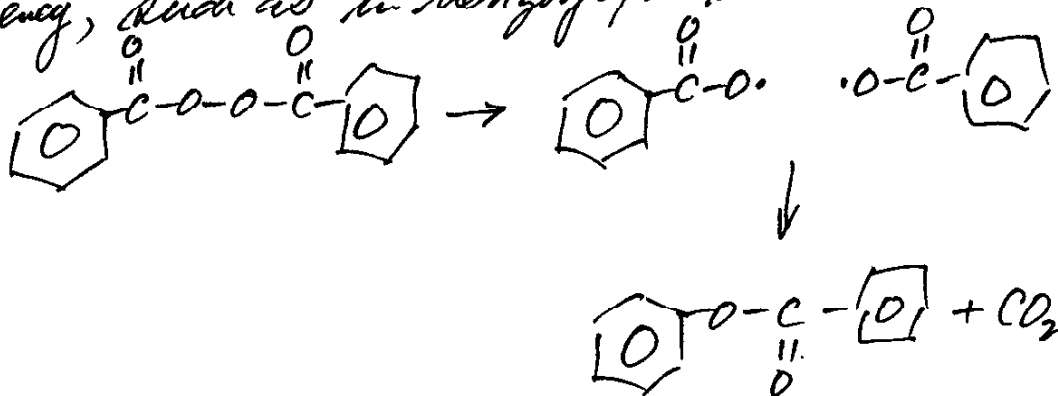
$$F. X_w/X_n = M_w/M_n = PDI$$

(4)

G. Free radical mechanism for termination, one chain gets reduced while its collision partner gets oxidized,

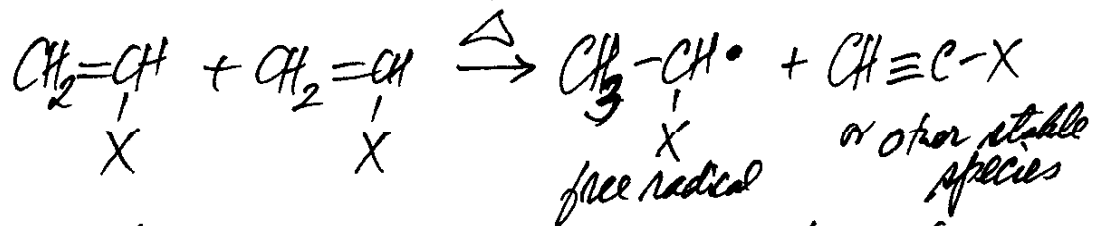


H. Free radical initiator fragments recombine within such a cage, resulting in a reduced initiator efficiency, such as in benzoyl peroxide



within a solvent cage

I. Some monomers can self initiate to start polymerization



J. Mercaptan or chlorohydrocarbon can act as the chain transfer agents



(5)

(A)

A.

- a) i) no loop forming side reaction
 ii) principle of equal reactivity, such that

$$\frac{C_0 - C_t}{C_0} = p \text{ which is indep of } x, \text{ chain length}$$

- iii) When $p \lesssim 1$, $q \equiv 1 - p \gtrsim 0$, such that $1 - p \approx e^{-q}$

- b) i) Steady-state approximation for

$$\frac{d}{dt} \sum_{x=1}^{\infty} (RM_x^\bullet) = 0 \quad \text{and} \quad \frac{d(RM_x^\bullet)}{dt} = 0$$

$$\text{such that } (RM_x^\bullet) = (R^\bullet) \alpha^{x-1}$$

$$\alpha = \frac{k_{p(M)}}{k_{p(M)} + (\gamma_i k_t)^{1/2}} \lesssim 1$$

- ii) $RM_x^\bullet + RM_y^\bullet \longrightarrow RP_{x+y}R$, recombination only

- iii) $1 - \alpha \equiv q \gtrsim 0$, such that $1 - \alpha \approx e^{-q}$

- c) i) no initiation or termination such that all reactions have the same rate constant k

- ii) $x \approx v$ where all $X(x)$ is populated such that

$$X_x = \frac{e^{-v} v^{x-1}}{(x-1)!} \approx \frac{e^{-(x-v)^2/2v}}{\sqrt{2\pi v}}$$

B.

$$W(x) = \frac{x X(x)}{\int_0^{\infty} x X(x) dx} = \frac{x X(x)}{Q_1}, \quad Q_1/Q_0 = x_m, \quad Q_0 = 1$$

$$\therefore x_m = Q_1$$

$$a) \quad W(x) = \frac{x q e^{-qx}}{(1/q)} = x q^2 e^{-qx}$$

$$b) \quad W(x) = \frac{x^2 q^2 e^{-qx}}{(2/q)} = \frac{1}{2} x^2 q^3 e^{-qx}$$

$$c) \quad W(x) = x e^{-(x-v)^2/2v} / \sqrt{2\pi v} \cdot v = \frac{x e^{-(x-v)^2/2v}}{\sqrt{2\pi v^3}}$$

(6)

$$a) \chi_w = \int_0^{\infty} \chi W(\chi) d\chi = q^2 \int_0^{\infty} \chi^2 e^{-q\chi} d\chi = q^2 \frac{2!}{q^3} = \frac{2}{q}$$

$$\chi_w/\chi_n = 2/q / 1/q = 2$$

$$b) \chi_w = \frac{1}{2} q^3 \int_0^{\infty} \chi^3 e^{-q\chi} d\chi = \frac{1}{2} q^3 \frac{3!}{q^4} = \frac{3}{q}$$

$$\chi_w/\chi_n = 3/q / 2/q = 3/2 = 1.5$$

$$c) \chi_w = \frac{1}{\sqrt{2\pi}v^3} \int_0^{\infty} \chi^2 e^{-(\chi-v)^2/2v} d\chi = \frac{1}{\sqrt{2\pi}v^3} \int_{-\infty}^{\infty} (u+v)^2 e^{-u^2/2v} du$$

$$\chi - v = u$$

$$d\chi = du$$

$$\chi=0, u=-v \rightarrow -\infty$$

$$\chi=\infty, u=\infty$$

$$= \frac{2}{\sqrt{2\pi}v^3} \left\{ \frac{1}{2^2(\frac{1}{2v})} \sqrt{\frac{\pi}{\frac{1}{2v}}} + v^2 \frac{1}{2} \sqrt{\frac{\pi}{\frac{1}{2v}}} \right\}$$

$$= \frac{1}{\sqrt{2\pi}v} \left\{ v\sqrt{2\pi} + v^2\sqrt{2\pi} \right\} = 1+v$$

$$\chi_w/\chi_n = \frac{v+1}{v} = 1 + \frac{1}{v} \approx 1$$

C. $W(\chi)$ maximizes when $\chi = \chi_n$

$$a) \frac{dW(\chi)}{d\chi} = q^2 (e^{-q\chi} + \chi(-q)e^{-q\chi}) = 0$$

$$1 = q\chi_{max}, \quad \chi_{max} = \frac{1}{q} = \chi_n$$

$$b) \frac{dW(\chi)}{d\chi} = \frac{1}{2} q^3 (2\chi e^{-q\chi} + \chi^2(-q)e^{-q\chi}) = 0$$

$$2 = q\chi_{max}, \quad \chi_{max} = \frac{2}{q} = \chi_n$$

$$c) e^{-(\chi-v)^2/2v} \rightarrow e^0 \text{ when } \chi=v=\chi_n$$

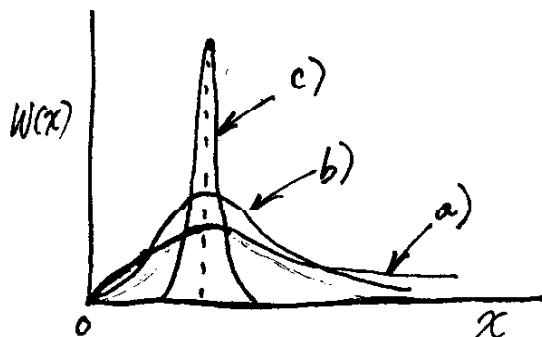
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(4) D. a) $W(\frac{1}{q}) = \frac{1}{q} \cdot q^2 \cdot e^{-q \frac{1}{q}} = q e^{-1} = q/e$

b) $W(\frac{2}{q}) = \frac{1}{2} (\frac{2}{q})^2 q^3 e^{-q \frac{2}{q}} = \frac{1}{2} \frac{4}{q^2} \cdot q^3 \cdot e^{-2} = 2q/e^2$

c) $W(v) = \frac{v}{\sqrt{2\pi} \cdot v} e^0 = \frac{1}{\sqrt{2\pi}}$

E.



a) $W(\frac{1}{q}) = \frac{1}{10^3 \cdot e} = 3.7 \cdot 10^{-4}$

b) $W(\frac{2}{q}) = \frac{4}{10^3 \cdot e} = 1.5 \cdot 10^{-3}$

c) $W(v) = \frac{1}{\sqrt{2\pi} \cdot 10^3} = 1.3 \cdot 10^{-2}$

Not necessary

(5) A. $F_1 = \frac{-d(M_1)/dt}{-\frac{d(M_1)}{dt} - \frac{d(M_2)}{dt}} = \frac{d(M_1)}{d(M_1) + d(M_2)}$

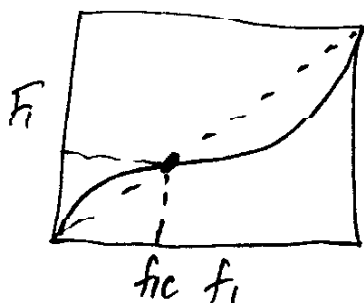
B. Steady state approximation for (R_1) and (R_2)

$-\frac{d(R_1)}{dt} = k_{12}(R_1)(M_2) - k_{21}(R_2)(M_1) = 0$

$-\frac{d(R_2)}{dt} = -k_{12}(R_1)(M_2) + k_{21}(R_2)(M_1) = 0$

$\therefore k_{12}(R_1)(M_2) = k_{21}(R_2)(M_1)$

C



f_{1c} : azeotropic composition

$f_{1c} = \frac{n_1 f_{1c}^2 + f_{1c}(1-f_{1c})}{n_1 f_{1c}^2 + 2f_{1c}(1-f_{1c}) + n_2(1-f_{1c})^2}$

$f_{1c} = \frac{1-n_2}{2-(n_1+n_2)}$

This should suffice

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D. $f_{ic} = \frac{1}{3}$

$$\frac{1}{3} = \frac{r_1 \left(\frac{1}{3}\right)^2 + \left(\frac{1}{3}\right)\left(\frac{2}{3}\right)}{r_1 \left(\frac{1}{3}\right)^2 + 2\left(\frac{1}{3}\right)\left(\frac{2}{3}\right) + r_2 \left(\frac{2}{3}\right)^2} = \frac{r_1 \frac{1}{9} + \frac{2}{9}}{r_1 \frac{1}{9} + \frac{4}{9} + r_2 \frac{4}{9}} = \frac{r_1 + 2}{r_1 + 4 + 4r_2}$$

$$3r_1 + 6 = r_1 + 4 + 4r_2, \quad 2r_1 + 2 = 4r_2, \quad \underline{\underline{r_1 + 1 = 2r_2}}$$

E. 1

